

## Equilibrium And Pressure Gizmo Answers

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Equilibrium and Concentration. Observe how reactants and products interact in reversible reactions. The initial amount of each substance can be manipulated, as well as the pressure on the chamber. The amounts, concentrations, and partial pressures of each reactant and product can be tracked over time as the reaction proceeds toward equilibrium.

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This is an example of Le Châtelier's principle, which is the tendency of a system in equilibrium to shift in response to a change. Apply: Select Reaction 5. If you were making water ( $H_2O$ ) from oxygen ( $O_2$ ) and hydrogen ( $H_2$ ), would it be an advantage to increase the pressure? Explain your

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answer, and then use the Gizmo to check your work.

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Name: \_\_\_\_\_ Date: \_\_\_\_\_ Student Exploration: Equilibrium and Concentration Vocabulary: chemical equilibrium, concentration, equilibrium, equilibrium constant, reaction quotient, reversible reaction, Le Châtelier's principle Gizmo Warm-up If Gary spends exactly as much as he earns, his savings will be in equilibrium.

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At a particular temperature,  $K_p = 0.240$  for the reaction  $N_2O_4(g) \rightleftharpoons 2NO_2(g)$  A flask containing only  $N_2O_4(g)$  at an initial pressure of 4.50 atm is allowed to reach equilibrium. Calculate the total pressure in this flask at equilibrium. With no change in the amount of material in the flask, the volume of the container in question is decreased to 0.400 times the original volume.

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Student Exploration: Equilibrium and Pressure. Vocabulary: Dalton's law, Le Châtelier's principle, partial pressure, pressure. Prior Knowledge Questions (Do these BEFORE using the Gizmo.) A typical scuba tank has a volume of 11 liters and can support a diver for one hour. An adult breathes about 3 liters of air with each breath.

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Partial pressure and equilibrium Get the Gizmo ready: Select Reaction 2. Move the Sim. speed slider all the way to the right. Introduction: In the Equilibrium and Concentration Gizmo, you found that reversible reactions eventually result in chemical equilibrium. Chemical equilibrium is reached when the rates of the

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Equilibrium and Pressure. Observe how reactants and products interact in reversible reactions. The amounts of each substance can be manipulated, as well as the pressure on the chamber. This lesson focuses on partial pressures, Dalton's law, and Le Chatelier's principle.

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