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Wanted to Know.

CHEMISTRY CHAPTER 2

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Measurements and

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Calculations SECTION 1
SHORT ANSWER

Answer the following questions in the space provided. 1. Determine whether each of the following is an example of observation and data, a theory, a hypothesis, a control, or a model.

observation and data
a. A research team records the rainfall in inches

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and Chemical
Compounds

$2\text{N}_2(\text{g}) \rightarrow 4\text{N}_2(\text{g})$ a.

Balance the equation.

$2\text{N}_2(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow$

$4\text{NO}_2(\text{g})$ 4 mol NO_2 :3

mol O_2 b. What is the
mole ratio of NO_2 to O_2 ?

15.0 mol c. If 20.0
mol of NO_2 form, how
many moles of O_2

must have been
consumed? True d.

Twice as many moles
of NO_2 form as moles
of N_2O are consumed.

True or False? False e.

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Twice as many grams
of NO₂ ...

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Chapter 2
Measurements And
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Table 2-5 page 47.
Sample problem 2-4
page 47. 2. Rounding.
Important when doing
calculations -
especially if using a

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calculator. Can't show more significant figures than necessary. Table 2-6 page 48 - skip last three rules. 3. Addition or Subtraction with Significant Figures

6 Chemical Bonding

CHAPTER 5 REVIEW

The Periodic Law

SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table,

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elements are ordered
(a) according to
decreasing atomic
mass. (b) according to
Mendeleev's original
design. (c) according to
increasing atomic
number. (d) based on
when they were
discovered. 2. d
Mendeleev noticed that
certain similarities in
the ...

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3.1 discusses some of the most fundamental assertions of chemistry, including the law of conservation of mass, the law of definite proportions, and the law of multiple proportions. It also gives a ...

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CALCULATIONS Include graphic organizer(s) for this chapter The following pages contain the bulk (but not all) of the information for the chapter 2 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test.

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CHAPTER 6 REVIEW

Chemical Bonding

SECTION 3 SHORT

ANSWER Answer the following questions in the space provided. 1.

a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c)

crystal. (b) molecule.

(d) atom. 2. d In a

crystal of an ionic compound, each cation is surrounded by a

number of (a)

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molecules. (c) dipoles.
(b...

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CHAPTER 3 REVIEW

Atoms: The Building

Blocks of Matter

SECTION 2 SHORT

ANSWER Answer the following questions in the space provided. 1.

In cathode-ray tubes, the cathode ray is emitted from the negative electrode,

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which is called the cathode. 2. The smallest unit of an element that can exist either alone or in molecules containing the

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CHAPTER 1 REVIEW

Matter and Change

MIXED REVIEW SHORT

ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a

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homogeneous or
heterogeneous
substance.

homogeneous a. sugar

homogeneous d.

plastic wrap

homogeneous b. iron

filings heterogeneous

e. cement sidewalk

heterogeneous c.

granola bar 2. For each

type of investigation,

select the most

appropriate ...

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Polymers: Chemistry and Physics of Modern Materials, Third Edition
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The Measurements and Calculations chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with measurements and calculations.

1 Matter and Change -
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Answers: 2

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Measurements and
Calculations SECTION
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SECTION 2-3 continued
4. Carry out the
following calculations.
Express each answer
to the correct number
of significant

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Answer Key ... Average
Customer Review: 3.4
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5 *The Periodic Law*

Answers SECTION 1 continued

6. Some binary compounds are ionic, others are covalent. The type of bond favored partially depends on the position of the elements in the periodic table.

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