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CHAPTER 6 REVIEW Chemical Bonding
SECTION 4 SHORT ANSWER Answer the
following questions in the space
provided. 1. _____ In metals, the valence
electrons are considered to be (a)
attached to particular positive ions. (c)

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immobile. (b) shared by all surrounding atoms. (d) involved in covalent bonds. 2.

~~7 Chemical Formulas and Chemical Compounds~~

CHAPTER 5 REVIEW The Periodic Law
SECTION 1 SHORT ANSWER Answer the
following questions in the space
provided. 1. c In the modern periodic

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table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered.

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Chapter: Chemical Bonding In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. _____ 1. The charge on an ion is a. always positive. b. always negative. c. either positive or negative. d. zero.

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~~16 Reaction Energy~~

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SECTION 1 continued 6. Some binary compounds are ionic, others are covalent. The type of bond favored partially depends on the position of the elements in the periodic table.

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CHAPTER 16 REVIEW Reaction Energy
SECTION 2 SHORT ANSWER Answer the following questions in the space

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provided. 1. For the following examples, state whether the change in entropy favors the forward or reverse reaction:
forward reaction a. $\text{HCl(l)} \rightarrow \leftarrow \text{HCl(g)}$
reverse reaction b. $\text{C}_6\text{H}_{12}\text{O}_6(\text{aq}) \leftrightarrow \text{C}_6\text{H}_{12}\text{O}_6(\text{s})$
forward reaction c. $2\text{NH}_3(\text{g})$
...

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Reviewed by Andrew Bonham, Professor,

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Metropolitan State University of Denver
on 7/16/19 Modern chemistry chapter 6
5 review answers. This textbook
addresses all the major topics that I
would expect to see in a General,
Organic, and biological Chemistry
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SECTION 2 SHORT ANSWER Answer the
following questions in the space
provided. 1. Use the concept of potential

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energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the

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